

Significant Figures in Calculations Worksheet
Chemistry

Name: _____

Date: _____ Hour: _____

List the number of significant figures in each of the following:

1) 127.34 g _____ 3) 120.007 μmol _____

2) 0.04060 cm _____ 4) 109 000 kA _____

Round each of the following measurements to TWO significant figures:

1b) 127.34 g _____ 3b) 120.007 μmol _____

2b) 0.04060 cm _____ 4b) 109 000 kA _____

Do the indicated calculation and write the answer with the appropriate number of significant figures:

5)
$$\begin{array}{r} 2.34 \text{ g} \\ 12.8 \text{ g} \\ + \underline{20.234 \text{ g}} \end{array}$$

6)
$$\begin{array}{r} 234 \text{ g} \\ + \underline{0.04 \text{ g}} \end{array}$$

7)
$$\begin{array}{r} 23.27 \text{ km} \\ - \underline{12.058 \text{ km}} \end{array}$$

8)
$$\begin{array}{r} 13.57 \text{ g} \\ - \underline{6.3 \text{ g}} \end{array}$$

9) $2.6 \text{ cm} \times 3.78 \text{ cm} =$ _____

10) $\frac{0.6717 \text{ g}}{3.99 \text{ cm}^3} =$ _____

11) $35 \text{ cm}^2 / 0.62 \text{ cm} =$ _____

12) $(0.08685 \text{ m})(28\,000 \text{ m}) =$ _____