

Here is a list of topics/concepts covered on September 27th's quiz – understand them, be able to explain them, and be able to give examples (and non-examples) of each.

chemistry (definition)	matter	energy
potential energy	kinetic energy	temperature
law of conservation of mass	qualitative data	quantitative data
law of conservation of energy	SI	7 base units of SI
Einstein and his famous equation	major prefixes of SI	mass
scientific law vs. scientific theory	weight	symbol v. abbreviation
accuracy	precision	percent error
important temperatures to know	three temperature scales	percent error formula

Practice questions:

- What are the seven base units of the SI system? _____

- What is chemistry? _____

- (a) What is the difference between an abbreviation and a symbol? _____

 - Give an example of an abbreviation and a symbol: _____
 - Why doesn't SI use abbreviations? _____

- Give three examples of possible QUALITATIVE data from our first lab.

- Give three examples of possible QUANTITATIVE data from our first lab.

- How is a scientific theory different from a scientific law? _____

- State the law of conservation of mass. _____

8. State the law of conservation of energy. _____

9. Explain what the letters in Einstein's equation stand for:

10. List two ways mass is different from weight: _____

11. A student measured the amount of solid produced by the reaction as 45.0 g. The literature said he should have made 42.0 g. What was his percent error? [7.1%]
12. The answer to the lab was 4.04 g. You got 3.12 g. What was your percent error? [23%]
13. What is accuracy? When is a set of measurements considered accurate? _____

14. What is precision? When is a set of measurements considered precise? _____

15. How can you test if a balance is accurate? _____

16. How can you test if a balance is precise? _____

17. What is temperature related to? _____
18. What does SI stand for? _____