

Calculations:

- 1) mass of empty flask = \_\_\_\_\_ g
- 2) mass of magnesium ribbon = \_\_\_\_\_ g [show your work!]
- 3) mass of hydrochloric acid = \_\_\_\_\_ g [show your work!]
- 4) total mass before reaction = \_\_\_\_\_ g [show your work!]  
(1 + 2 + 3)
- 5) mass after reaction = \_\_\_\_\_ g
- 6) mass of hydrogen gas = \_\_\_\_\_ g [show your work!]  
(4 - 5, should be positive)
- 7) ratio of mass of hydrogen gas to mass of magnesium metal = [show your work!]  
 $\frac{6}{2} =$  \_\_\_\_\_ (no unit)

**\*\*THIS IS THE ANSWER TO THE LAB'S GOAL!!\*\***

- 8) Write a sentence about your precision after gathering the [answers from three other groups](#).
- 9) Calculate your accuracy by calculating your percent error. (NOTE: The correct answer is on the [original lab handout!](#))