

$K_{sp}$  and Precipitates Practice  
Chemistry II

2 points

Name: \_\_\_\_\_

Date: \_\_\_\_\_ Hour: \_\_\_\_\_

Two solutions, 150.0 mL of  $1.00 \times 10^{-2}$  M magnesium nitrate and 250.0 mL of  $1.00 \times 10^{-1}$  M sodium fluoride, are mixed together.

- A) Write and balance the full reaction equation. Include states of matter.
- B) Determine if there will be a precipitate. (Hint: Q!) If there is, write the net ionic equation.
- C) Calculate the equilibrium concentrations of magnesium ions and fluoride ions.